

Title : Science course unit 10
(Covalency)

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Roger Hill (R.R.Hill
on credits)

CU S100/10

Tape No. 6LT/10056

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Form VTR

1st TX: 13.3.71.

Seq.	Time	Footage	Sequence List	Sound Cue
1	35"		Prof. Pentz introduces the unit.	
2.	6'55"		R.R. Hill illustrates the shapes of covalent molecules with balloon analogues. He models the following: Boron Trifluoride Beryllium Dichloride Phosphorous Pentafluoride Sulphur Hexafluoride	541.224 for covalent molecules
3	8'25"		Prof. Pentz introduces the principle of chirality. In an experiment he superimposes a teacup on its own image.	you'll have seen
	10'00"		Pentz tries to do the same with an oil can. It cannot be superimposed on its image and is therefore chiral.	
	13'21"		Pentz introduces the plane of symmetry concept as a useful tool for determining chirality of objects and molecules. He then changes the chirality of a teacup by moving a black dot which was bisected by the plane of symmetry to another place on the cup.	a chiral one
4	14'55"		P.R. Hill explains the importance of chirality at the molecular level. He shows a model of a glucose molecule together with one of its mirror image. It cannot be superimposed on its image.	well, chirality is-----
	19'09"		(+) Glucose and its image (-) glucose are examined with a polarimeter. The polarised light is shifted in different directions by each type of glucose. Molecules that react in this way are called optical isomers.	547.12252 547.7813 541.7 541.7028
5			How do enzymes in a living organism distinguish between optical isomers? R.R. Hill demonstrates with an analogue. He places his hands (optical isomers) into a cloth bag (ordinary chemical reagent). Both hands fit. He then tries to place each hand into a glove (enzyme). The left hand will not fit the right hand glove.	547.758

